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"STUDY OF CHEMICAL KINETICS OF URANIUM MINING BY ACID LEACHING." "Исследование кинетики процесса добычи урана сернокислотными растворами."

Kazakhstan produce more than 40% of world Uranium and all of Uranium produced in Kazakhstan are mined by In-Situ leaching (ISL). The ISL involves recovering the minerals from ground by dissolving them and pumping the pregnant solution to the surface where the minerals can be recovered.

There are two different regimes of uranium leaching depending on ore mineralization and carbonization: acid, sulfuric acid or less commonly nitric acid, and carbonate. All of Kazakhstani uranium deposits are mined by acid leaching.

It is known, that mostly uranium ore contain tetra and hexavalent uranium oxide, and tetravalent uranium oxide hardly reacts with the acid. Solution of tetravalent uranium oxide by sulfuric acid requires oxidizers which transform tetravalent uranium oxide to hexavalent one. In present paper the influence of oxidants, such as Iron (III) Hydroxide and Manganese dioxide, on the Uranium recovery is investigated.

A model of the uranium mining process with detailed kinetics, taking into account the redox processes in the ore formation is proposed. Tetravalent uranium is oxidized by ferric hydroxides, and proceeds to the hexavalent state. However, this leads to a reduction in the formation of iron (III) hydroxide, which can recover manganese oxides. It was found that the presence of iron (III) hydroxide and manganese oxide in the formation significantly increases the recovery of uranium, which is consistent with practical results.